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Synthesis of aspirin lab report introduction

LAUC Aspirin Synthesis

Experiment
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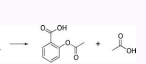
Aspirin is the common name for the compound acetylsalicylic acid, widely used as a fever reducer and as a pain killer. Salicylic acid, whose name comes from Salix, the willow family of plants, was derived from willow bark extracts. In folk medicine, willow bark teas were used as headache remedies and other tonics. Nowadays, salicylic acid is administered in the form of aspirin which is less irritating to the stomach than salicylic acid. To prepare aspirin, salicylic acid is reacted with an excess of acetic anhydride. A small amount of a strong acid is used as a catalyst which speeds up the reaction. In this experiment, phosphoric acid will be used as the catalyst. The excess acetic acid will be quenched with the addition of water. The aspirin product is not very soluble in water so the aspirin product will precipitate when water is added. The synthesis reaction of aspirin is shown below.



Since acetic acid is very soluble in water, it is easily separated from the aspirin product. The aspirin isolated in this step is the "crude product". A "purified product" can be obtained through recrystallization of the crude product in hot ethanol. In this experiment, the crude product will be the desired product. The percent yield of the crude product will be determined for this reaction. The purity of the product will also be analyzed. The product will be analyzed by three different methods: melting point, titration, and spectroscopic assay.

The melting point range of pure aspirin is 135-140 °C and the melting point range of the salicylic acid starting material is 158-161 °C. If impurities are present in your crude sample, the melting point range for your product will be lower than the range of pure aspirin. Also, your melting point range may be greater than 2 degrees.

From the titration of your sample, the moles of acetylsalicylic acid present can be determined assuming that there is not a large percentage of an acid impurity present in your crude sample.



LAB REPORT EXPERIMENT 21
Synthesis of Aspirin (Acetylsalicylic Acid)

I. Crude Product

Reagent	Molecular Weight	Grams	Moles
Salicylic acid		42.5 g	
Acetic anhydride			

The yield limiting reagent is _____ mol

Theoretical yield of aspirin: _____ g

Actual yield: _____ g

Percentage yield: _____ %

Observed melting point of aspirin, **crude product**: ~~130-135~~ 139°C

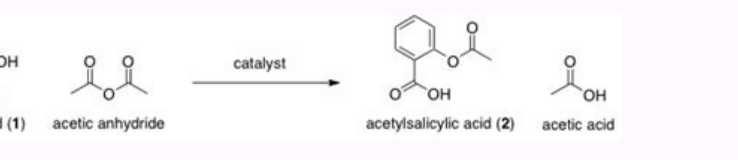
Literature melting point of aspirin: 136°C

Results of the ferric chloride test: positive

Test tube 1: insoluble (purple)

Test tube 2: soluble (purple)

Test tube 3: two layers, purple, soluble



Post-Laboratory Report

Experiment: Synthesis of Acetylsalicylic Acid (Aspirin)

Summary: The synthesis of aspirin by reacting salicylic acid with acetic anhydride in the presence of phosphoric acid as a catalyst to yield acetylsalicylic acid and acetic acid. The reaction was carried out in a round-bottom flask equipped with a magnetic stirrer. The reaction mixture was heated in a water bath at 60°C for 15 minutes. The excess acetic acid was quenched with water, and the aspirin product was isolated by filtration and recrystallization. The crude product was obtained in a yield of 76.7%.

Chemical Reaction:

Salicylic acid

Acetic anhydride

Acetylsalicylic acid
(Aspirin)

Experimental Data:

Name	Initial Mass (g)	Final Mass (g)	Melting Point (°C)	Boiling Point (°C)	Yield (g/mol)
Salicylic acid	42.5	42.5	136	136	138
Acetic anhydride	102	102	102	102	102
Phosphoric acid	10	10	10	10	10
Crude acetylsalicylic acid	180	180	139	139	180

The reaction in which the carboxylic acid $\text{C}_7\text{H}_6\text{O}_3$ is combined with an alcohol in the presence of a catalyst (commonly concentrated sulfuric acid) to form a $\text{C}_9\text{H}_8\text{O}_4$ is called esterification reaction. Once made water is added from heat with cold water addition and suction filtration is carried out. Since the values are quite close, this shows that the aspirin we obtained is quite pure and, therefore, it contained fewer impurities. If it is separated solid, calcium until the solid is completely dissolved. The solution was allowed to cool. The dry glass is weighted together with the filter paper and the clock glass. As for the recrystallization of aspirin, the gross product collected prepared in the preparation of the aspirin that dissolves in ethanol is added and hot distilled water is added to the solution. Therefore, this could have affected the percentage performance as well. Share this: Facebook Twitter Reddit LinkedIn WhatsApp A, is a reversible reaction and products of sweet odor products. It is used in the polymer industry. Once solids dissolves, weigh the glass paper and filter paper, use the filter paper to carry out the suction filtration, place crystals in the glass of the clock, spot the dry crystal and calculate the weight of the aspirin. 3 to 4 drops of conc. The theoretical merger temperature is $140 \text{ }^\circ\text{C}$. In this experiment, the percentage performance is 76.7% performance

that shows that it is almost completely pure and contains minor impurities. Therefore, we had to continue filtering the entire mixture into the filter bottle and, therefore, this could be led to the presence of more impurities or loss in the product. Aspirin theory is a medication that is generally used to alleviate minor pain and pain and other medical uses, such as anti-inflammatory medications. From both calculations, Evaluate that aspirin is pure to a large extent, however, due to some incorrect errors or techniques. The percentage of performance is not high and incomplete reactions can also be one of the reasons. 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If the solid is separated, it dissolves completely. Procedure The preparation of aspirin 2.4 g of salicylic acid is measured in a and it's logged. However, there will be some environmental effects that will still affect the experiment to a small extent. Dry glass is weighed together with filter paper and watch glass. As we have to wash the crystals before carrying out suction filtration, some crystals may have dissolved. An additional 40mL of cold water is added. The melting point is determined using the optimelt melting apparatus. It studies the identification and manipulation of age-dependent modifiers of cardiac function and mutation, hierarchical modeling and imaging of contractile machinery, integrative analysis of striated muscle performance and myopathic processes. The percentage yield was about 76.7% while the temperature range is between 134.2 and 136.1 ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ ºÁ It measures the weight of dry and recrystallized aspirin, the expected yield of aspirin from the amount of salicylic acid, the percentage yield of dry and recrystallized aspirin. However, there is also another possibility for the yield value of the lowest percentage. The main experimental findings are that precision and attention is very important in this experiment to get aspirin that is pure. The synthesis of aspirin from sodium salicylate and acetic anhydride, followed by analysis by infrared spectroscopy, determination of the melting point and thin layer chromatography5PY022 Áç Á Á Á Á Á Á Á Á Á Á Á Á Á y Usage Marquis GonzalezSynthesis of AspirinIntroduction:The aim of this experiment is to allow us to perform the synthesis of aspirin, strengthen the recrystallization skills and strengthen the technique of determining the melting point. The mixture is heated for 10 to 15 min to complete the reactions.The filtration is done by suction and the raw product collected after washing it a little with cold water.An additional 40ml of Friar. The mixture is heated for 10 to 15 min to complete the reactions. This could have affected the results. 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Aspirin is a derivative of salic acid which is a soft analog and not narcÁ º useful for relieving headache and muscle and joint pains. The sulfuric acid is added to the mixture and stirred. Therefore, from this I learn that I must be more alert when I perform º suction leak º avoid unnecessary discomfort and inaccuracies of results and calculation. We investigate conserved myopathic processes (muscle disease) and perform hierarchical and integrative analysis of muscular function º from the level of individual molecules and macromolecular complexes to the level of the tissue itself. The anhidrido acÁ© is used because it is cheap and forms a by-product, the acÁ© acid that is not corrosive and can be recovered to produce more anhidrido acÁ© unlike other acetylating agents that can also be used. The rich melting point º 140°C and the experimental melting point range is from 134.2°C to 136.1°C. The º is filtered by succiÁ º n and the raw product is collected after washing a little with water fraA. Water is removed once the heat is removed with º of water fraA and the filtration º by succiÁ º n is carried out. Therefore, the amount of water we use to wash the crystals during succiºn filtration³ may have affected the percentage yield as well. For the preparation º aspirin, the measured amount of salic acid is anhdrido ac©. The temperature is fixed at 100 C and then gradually heated to determine the point of fusiÁ º n. Experimental errors There were some experimental errors that have caused variations in my results compared º with the º rich solutions. Recrystallization º Aspirin The crude product collected was dissolved in approximately 5 ml of ethanol in a flask 100 ml. It is also useful in which unique orgÁ © for testing alcohols and acid carboxÁlco. Conclusion³ From this experiment, I have learned how º carry out the succiÁ º n succiÁ º succiÁ º , ³simehc namreG a ,namffoH xileF noitcudortnl .rebur citehtnys dna spaos, seyð, senicidem, sreuqcal, sehsnrav, stniap fo erutcafunam eht ni desu si noitcaer noitacifiretsE .spets noitacifirup dna noitalosi eht gnirud gnirrucco netfo tucudorp fo ssol si ereht esuaceb si egatnecrep laciteroeh eht morf tereffid si egatnecrep dley latnemirepxe ehT .nirpsa eht fo tniop gnitlem eht gninimreted yltsal dna nirpsa fo noitasillatsyrceR ,nirpsa fo noitaraperp era serudecorp niam ehT .etarucca etiuq si ti sa elbatpecca erew stluser eht dna tem si tnemirepxe eht fo evitcejbo eht yllaniF .lohocla na dna dica na morf retse mrof ot tuo deirrac si tI .snoitcaer derisednu morf tnemirepxe dna tucudorp eht toctorP tnemirepxe eht tuo gniyrac resu eht toctorP: swollof sa era doohemuf eht fo esoprup niam eerht ehtT .tnemtrapeD ygoloidraC eht ni enicidem fo rosseforp tnatsissa na si, Dm, otarammac ssor ynohnta .noitanimreted tniop gnitlem fo euqinhcet eht ecrofnier dna noitasillatsyrceR fo slliks eht e crofnier, nirpsa fo sisehtnys eht tcudnoc ot su elbane ot si tnemirepxe siht fo evitcejbo eht Á Á ppAstahW nidekniL tiddeR rettiwT koobecaF: siht erahS eliforf bal tidE baL hcraeseR a dñiF ygoloiisyhp elcsum, yhtapoym, noitcnuF elcsum, ygoloib elcsum detairts, sessecorp cihtapoym, sciteneG, tnempoleved elcsum: saerÁ hcraeseR .noitacifiretse sa nwonk si nirpsa fo sisehtnys stluser ni ycarucca erpuse ot ssecorp siht tuo gniyrac nehW morf yawa tpek eb dna fo dehctiw eb tsuim ecruos dnw lia taht trael dah l, erofereñI .cnoc fo sport 4 ot 3.doohemuf eht ni ksalf eht ot dica cilycilas eht ot dedda si edirrhna citeca fo Lm6.dedrocer si dna ksalf lacinoc lm001 a ni derusaem si dica cilycilas fo gq.2: doñeM edirdyhna citeca sseceæ eht esompoced ot toh si ti elihw erutum eht OT DEDDA YLSUOITUAC SI RETAW DELLITSID FO LM 1 , TAEH MORF GNIVOMER RETFA .Demraw .no OS DNA Slacimehc Gnildna Nehw Emit Lla Ta Suoituae EB OT DNA YAW THGIR EHT a stable form of acetylsalic acid, commonly known as aspirin, in 1997. A smoke bell is a large piece of equipment with five closed sides of a work area, the bottom of which is usually at a standing working height. We are interested in the basic mechanisms of striated muscular biology. First, after obtaining the raw product º the first suction filtration³ we had to transfer it to the single flask³ to perform the recrystallization³ We use a variety of imaging techniques to study the “structural physiology” of the skeletal and caracal muscle. Principal Investigator: Anthony Cammarato, Ph.D. Department: The Cammarato Laboratory is located in the Division º Cardiology of the Department of Medicine of the School of Medicine of Johns Hopkins University. The temperature range of aspirin fusiÁ º according to my experiment is between 134.2 Á C and 136.1 Á C.

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